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Enclosed herewith for filing is a patent application, as follows:

Inventor(s): Zhang, Fan; Towery, Daniel L.; Levert, Joseph A.  
Title: **Composition for Chemical Mechanical Planarization of Copper, Tantalum and Tantalum Nitride**

- ☒ Return Receipt Postcard
- ☒ This Transmittal Letter (in duplicate)
- 3 pages Declaration For Patent Application and Power of Attorney
- 26 pages Specification (not including claims)
- 3 pages Claims
- 1 page Abstract
- 1 Sheet of Drawing
- 1 page Recordation Form Cover Sheet
- 2 pages Assignment
- 1 page PTO Form 1449 citing 5 references

**CLAIMS AS FILED (fees computed under §1.9(f))**

For	Number Filed		Number Extra		Rate	\$	Basic Fee
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COMPOSITION FOR CHEMICAL MECHANICAL PLANARIZATION OF  
COPPER, TANTALUM AND TANTALUM NITRIDE5 BACKGROUND OF THE INVENTIONTechnical Field

This invention relates to chemical mechanical processes for the planarization of surfaces, and to chemical compositions especially suited thereto.

10 More particularly, this invention relates to composition for the chemical mechanical planarization of conductive and insulating surfaces as typically encountered in the fabrication of integrated circuits, including copper, tantalum and tantalum nitride.

15 Description of Related Art

Fabrication of integrated circuits ("ICs") to improve performance and reduce costs involves complex analysis of materials properties, processing technology and IC design. IC's consist of multiple layers of conducting, insulating and semiconductor materials, interconnected in various ways by conducting

20 metallic channels and plugs ("vias"), including various dopants implanted into various materials for producing the electronic functionality desired of the IC. The near-universal trend in the manufacture of integrated circuits is to increase the density of components fabricated onto a given area of wafer, increase the performance and reliability of the ICs, and to manufacture the ICs at lower cost

25 with less waste and fewer defective products generated by the manufacturing process. These goals lead to more stringent geometric and dimensional requirements in the manufacturing process. In particular, etching precise patterns into a layer is facilitated by the layer having a surface as nearly planar as feasible at

the start of the patterning process. For the common case of patterning by means of photolithography, a planar surface permits more precise location and dimensioning for focusing the incident radiation onto the surface to be etched than would be possible with a surface having deviations from planarity. Similar conclusions typically apply for electron beam or other means of etching. That is, deviations from planarity of the surface to be etched reduce the ability of the surface to support precisely positioned and precisely dimensioned etches. In the following description of the present invention we focus on the typical etching, planarization and photolithography processes as practiced in the manufacture of ICs. However, this is by way of illustration and not limitation, as those skilled in the art of etching will appreciate that the techniques of the present invention producing planar surfaces will have applicability in increasing the precision of etching by means other than photolithography. In addition, the present invention is not limited to the field of IC manufacture and may find applicability in other areas of technology requiring planar surfaces.

Chemical Mechanical Planarization ("CMP") has been successfully integrated into integrated circuit multilayer manufacturing processes to achieve highly planar surfaces as described in text books (for example, "Microchip Fabrication" by Peter Van Zant, 3rd Ed., 1997) and generally known in the art. We note that "CMP" is also used in the art to denote "Chemical Mechanical Polishing" as well as "Chemical Mechanical Planarization". We use CMP herein synonymously in either sense without distinction.

A typical CMP process is depicted schematically in Figure 1. During a CMP process, the wafer, 1, is typically held inside a rotating carrier and pressed onto a rotating pad, 2, under pressure, 6, while an abrasive slurry, 5, (typically containing particles of abrasive such as  $\text{SiO}_2$ ,  $\text{Al}_2\text{O}_3$ , and the like) flows between the wafer and the pad. The slurry, 5, will typically contain reagents for chemically etching the wafer, 1, leading to chemical as well as mechanical removal of material. Thus, in the typical practice of CMP, material removal is effected by a combination of chemical attack and mechanical abrasion.

Typically, the wafer, 1, will be caused to rotate as depicted by 3 in Figure 1, while the polishing pad will itself rotate (4 in Figure 1). Figure 1 depicts the polishing pad and wafer rotating in the same direction (for example, clockwise when viewed from above as in Figure 1). However, this is merely for purposes of illustration and counter-rotation of wafer and polishing pad is also practiced. In addition to the rotation of the wafer depicted by 3 in Figure 1, the wafer, 1, may be caused to oscillate in the plane of the surface being polished, substantially perpendicular to the direction of the applied force, 6 (This oscillatory motion is not depicted in Figure 1).

Increasing the speed and performance of ICs typically calls for increasing the density of components on the wafer and increasing the speed at which the IC performs its desired functions. Increasing component density typically requires decreasing the size of conducting channels and vias (or plugs). It is well known that decreasing the cross-section of a current-carrying conductor increases the electrical resistance of the conductor for the same material. Thus, decreasing component size on ICs increases electrical resistance, degrading performance and perhaps leading to unacceptable heating. This is one reason IC developers have been looking for conducting materials for use in IC fabrication having lower electrical resistance. Present IC technology typically makes use of tungsten (chemical symbol W) and aluminum (Al) as conductors. Both have adequate electrical conductivities in present devices, but future generations of IC devices will preferably make use of yet higher conductivity materials. Copper (Cu) is among the leading candidates.

Increasing the density of IC components on the wafer also increases the capacitance of the circuits. That is, bringing charge carrying circuit elements closer together increases the capacitive coupling between such circuit elements. Higher capacitance is detrimental to circuit performance, especially for higher frequency operation as would typically be encountered in telecommunication applications and elsewhere. However, capacitive coupling between proximate circuit elements can be reduced by reducing the dielectric constant of the insulator

or insulating material separating the coupled circuit elements. Thus, in addition to seeking conductors with higher conductivities, insulators with lower dielectric constant ("low k") are also being sought for use as insulating layers in ICs.

Current multi-layer IC fabrication typically makes use of tungsten (W) CMP processes at each successive circuit level. Typically, blanket films of W, Titanium (Ti) and Titanium Nitride (TiN) are deposited. The films are then typically polished, thereby removing material resulting in (for example) W vias or "plugs" which are inlaid, typically in an SiO<sub>2</sub> dielectric layer. The W plugs act as electrically conducting paths between the metal lines of adjacent layers of the IC. Typically, the metal lines connected by W vias will consists of alloys of Al and Cu in present ICs. In typical present IC designs, Ti and TiN layers are used as barrier layers (to hinder unwanted diffusive intermixing of components during fabrication) and adhesion layers (to promote good bonding between otherwise poorly bound layers and avoid delamination). Such barrier and adhesion layers must also be removed during W CMP to reveal the inlaid W plugs. Desirable CMP for such ICs should remove the various layers equally and, thus, planarize in one CMP step. Commercially available W slurries can achieve almost the ideal 1:1 removal rate selectivity between W and Ti/TiN layers. This results in a very flat surface of the device wafers after W CMP. Thus, if other combinations of conductor and insulators are to be used in IC fabrication, adequate chemical reagents for CMP must be employed and adequate planarization processes must be used. Such are among the objects of the present invention.

The use of W vias, Al-Cu lines, and SiO<sub>2</sub> dielectric layers, although quite successful in present ICs, has inherent drawbacks that hinder attaining the circuit performance desired in future devices. The Al-Cu alloys and W vias are conductive, but less so than Cu. The SiO<sub>2</sub> dielectric layers (although good electrical insulators) have a relatively high dielectric constant, leading to deleterious capacitive effects. ("High k" typically denotes dielectric constants in the range of approximately 3 to 9.) The combination of relatively high resistivity metals and relatively high dielectric constant insulators reduces circuit speed and

reliability, particularly as the device geometry is reduced in future ICs below approximately 0.25  $\mu\text{m}$ , (that is 0.25 microns).

5 Metallic copper (Cu) has a lower resistivity than W or Al-Cu alloys. Therefore, Cu is becoming a popular choice for the interconnect metal to be used in future generation ICs. It is further envisioned that low k dielectric materials (that is, materials with dielectric constants, k, less than about 3.0) will be used in conjunction with Cu metallization to reduce capacitive effects. However, both of these choices bring accompanying challenges in the fabrication of high performance, low cost, reliable ICs. Low k dielectrics are often mechanically weak relative to conventional dielectrics and tend to delaminate under the stress of CMP, especially if the applied pressure, 6, in Figure 1 must be rather large in order to achieve adequate material removal rates. An adequate rate of material removal is required in order to achieve planarization in an acceptable period of time. Addressing these challenges, the focus of the present invention is on the use of copper, on barrier layers to avoid harmful diffusion of Cu, and slurry compositions for effective Cu CMP in the presence of effective barrier and adhesion layers.

15 In order to increase performance and reduce manufacturing costs, it is envisioned that Cu metal will most likely be used in future ICs in fabricating the metallic conducting channels within a layer and in the vias which connect adjacent layers. This will likely be accomplished using a "metal damascene" manufacturing approach. Damascene processing typically proceeds by depositing a blanket layer of metal on top of a patterned insulating or dielectric layer, thereby filling channels and vias in the patterned insulating layer. The blanket metal layer is then typically removed by CMP or etching revealing the inlaid metal channels and vias. It is envisioned that the metal of choice for the next generations of ICs will be copper. Therefore, to be concrete in our description, we will describe the practice of the present invention in connection with copper damascene processing. However, the present invention is not inherently so limited and other embodiments will be obvious to those having ordinary skills in the art.

Copper has the advantage of higher conductivity, but suffers from several complications which heretofore have delayed its adoption in ICs. Among copper's disadvantages is the fact that it is a very diffusive contaminant. That is, copper diffuses widely and easily through other materials typically used in the fabrication of ICs, seriously degrading electronic performance by doing so. It is among the objects of the present invention to address, eliminate or ameliorate some of these attendant drawbacks in the use of Cu metallization in the fabrication of ICs.

In addition to its high rate of diffusion, reaction products of copper with typical etching reagents have often resulted in non-volatile (or insoluble) reaction products. Thus, etching of Cu with convention CMP slurries has been difficult. Identification of a group of effective copper etching reagents having volatile or soluble reaction products (while maintaining adequate removal rate and selectivity) is among the objectives of the present invention.

Tantalum (Ta) and Tantalum Nitride (TaN) have been identified as promising barrier layer, or "liner metals", that will prevent harmful Cu diffusion. Because CMP is presently the most effective and well understood planarization technique, it is the natural method with which to undertake the planarization of Cu, Ta or TaN. Such Cu/Ta/TaN CMP requires slurries with high Cu and Ta/TaN removal rates and close to a 1:1 removal selectivity between Cu and the liner metals. However, Ta and TaN are mechanically hard and they do not react readily with most etching chemicals. For these reasons, CMP slurries having appropriate chemical formulations to obtain 1:1 selectivity between Cu and the liner metals have been difficult to achieve. Hence, at present there is no slurry commercially available for Cu CMP.

Typical experimental Cu slurries are composed of  $H_2O_2$ , various oxidizers, alumina and/or silica abrasive, and other chemical components, typically in acidic (low pH) solutions. (Tytgat *et. al.* 4,981,553; 5,098,571). These formulations typically give good Cu removal rate, but often achieve very low Ta/TaN removal rates, even when high polishing pressures are employed. Currently there are two common experimental approaches being employed for Cu CMP, both of which

suffer from disadvantages. In one approach (Brusic, "A Cautious Approach to the Removal of Ta in the CMP Polishing of Cu/Ta Structures, 193<sup>rd</sup> Electrochemical Society Meeting, May 1998) Cu CMP is conducted by using a two-step polishing process to remove Cu and Ta/TaN. The Cu and Ta/TaN layers are removed  
5 separately in sequence using two distinct slurries. This two step approach significantly complicates the fabrication processes and increases the cost of the CMP process applied to Cu. A single-step Cu CMP would be preferable, but it would require a slurry with 1:1 selectivity for Cu and Ta/TaN. One possible way to increase the removal rate of Ta/TaN layers is to dramatically increase the  
10 polishing downforce. However, a higher polishing downforce is contraindicated since it could damage the underlying low k materials, which are often mechanically weak and subject to delamination. Achieving a slurry with the required near 1:1 selectivity without the application of large polishing downforce is among the objects of the present invention.

15 Slurry formulations that react chemically with inert liner metals to achieve adequate removal rates and selectivity would be an improvement in Cu CMP. A single-step Cu CMP employing a slurry that provide 1:1 selectivity and high material removal rates at low polishing pressures, is highly desirable. However, a two-step CMP slurry that did not require high polishing pressures, though less  
20 desirable than the single-step slurry, would still be an improvement in the present art.

The present invention is described for the specific example of CMP slurries for Cu/Ta/TaN on IC wafers as this specific case is expected to be a primary area of applicability of the present invention. However, the compositions and processes  
25 of the present invention are not inherently limited to these particular instances. The present invention could be useful for processing many different types of metallic, dielectric, or organic layers, or mixtures and/or composites thereof, on numerous types substrate for numerous technical applications, as would be known to those skilled in the art.

BRIEF SUMMARY OF THE INVENTION

Chemical mechanical planarization of surfaces of copper, tantalum and  
5 tantalum nitride is accomplished by means of the chemical formulations of the  
present invention. The chemical formulations may optionally include abrasive  
particles and which may be chemically reactive or inert. Contact or noncontact  
CMP may be performed with the present chemical formulations. Substantially 1:1  
removal rate selectivity for Cu and Ta/TaN is achieved. In addition to 1:1 removal  
10 rate selectivity, the present invention provides adequate material removal rates  
without excessive downforce being necessary on (often delicate) low dielectric  
components.

BRIEF DESCRIPTION OF THE SEVERAL VIEWS OF THE INVENTION

- 5 Figure 1: Schematic depiction of Chemical Mechanical Planarization (“CMP”).

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DETAILED DESCRIPTION OF THE INVENTION

In the following description and figures, similar reference numbers are used to identify similar elements.

- 5           The Cu/Ta/TaN CMP slurries, or "etchants", described herein are improved formulations of chemicals typically used for CMP including abrasive slurries, metal etchants and cleaners, chemical polishes, brighteners, and pickling solutions, etc. Such chemical mixtures are typically comprised of one or more of the following
- 10   constituents:

TABLE A

- Abrasive Particles
- Mineral Acids
- 15   • Organic Acids
- Strong Bases
- Mineral Salts
- Organic Salts
- pH buffers
- 20   • Oxidizing Agents
- Organic and Inorganic Peroxides
- Corrosion Inhibitors
- Chelating Agents
- Surface Modifying Agents
- 25   • Liquid polymers
- Surfactants
- Solution Stabilizers
- Solvents (including water)

In general, CMP using the etchants described herein require a method of introducing the etchant onto the wafer surface and preferably a method for transporting the etchant across the wafer surface. The CMP chemical formulations of the present invention need not include abrasives in the etchant mixture and do not necessarily require mechanical contact between the wafer surface and another surface or body, such as a polishing pad. Thus, the etchants of the present invention may be used to good effect in both contact and noncontact CMP.

In addition to adjusting the type and concentration of etchant constituents, heating the etchant solution may be used to control the removal and rate of removal of material from the surface of the wafer. Temperature control of the reaction may be achieved by introducing the etchant onto the wafer preheated to the desired temperature, by heating the polishing pad other body and/or by maintaining the polishing at a constant temperature.

Improved CMP is achieved by means of etching solutions as described below. Abrasive particles are optionally included and may optionally react with the surface of the wafer being etched as well as provide mechanical removal of wafer material. Many suitable abrasive materials could be used, including conventional abrasives ( $\text{SiO}_2$ ,  $\text{Al}_2\text{O}_3$  and the like), and various nonconventional abrasives that are comprised of metals, solid elemental particles (for example carbon), polymer particles, oxides, carbides, fluorides, carbonates, borides, nitrides, or hydroxides of several metals, including, but not limited to, Al, Ag, Au, Ca, Ce, Cr, Cu, Fe, Gd, Ge, La, In, Hf, Mn, Mg, Ni, Nd, Pb, Pt, P, Sb, Sc, Sn, Tb, Ti, Ta, Th, Y, W, Zn, Zr, or mixtures thereof. These particles may be coated with a thin layer of another material, including but not limited to those described above. The potential advantages of the use of coated particles are expected to include decreasing cost by coating a less dense, inactive and inexpensive particle, such as  $\text{SiO}_2$ , with a chemically active, and often more dense and expensive active material such as  $\text{CeO}_2$ . The effective density of such particles will be less than solid particles comprising all chemically reactive material, and thus more stable in terms of particle settling according to Stokes Law which predicts a larger settling

velocity for particles having a higher density. Similarly for a given wt% of solids, slurries comprised of coated abrasive particles (typically less dense) will have a greater number of particles in a given volume of fluid and thereby present a greater abrasive surface area in contact with the wafer surface.

5 It is envisioned in the practice of the present invention that many of the particle systems described herein will be produced by means of the "sol" method. This typically involves growing the particles to their final size in solution. By growing the particles entirely in solution and remaining in solution for use (that is, never dried) there is no sintering or "necking" of the particles that will result in  
10 large agglomerate, which may be damaging to the sensitive IC layers, or underlying structures. Having avoided agglomeration, these particles are introduced into solvent systems very readily and at lower cost than conventional abrasives that typically must undergo additional and expensive particle size reduction and powder dispersion processing. The practice of the present invention  
15 makes use of several particle size distributions. A bi-modal particle size distribution, or a multi-modal particle size distribution, or a broad Gaussian particle size distribution, may all be employed in the practice of the present invention with typical particle sizes in the range 4 nm to 5 $\mu$ m. It is envisioned in the practice of the present invention that particle sizes greater than approximately  
20 5 $\mu$ m will not give satisfactory results, particularly in terms of increasing removal rates and reducing defects and scratches.

We note elsewhere herein typical components of the etching reagents useful in the practice of the present invention. Practical industrial applications may also require the reagent mixture to contain other additives to inhibit premature reaction,  
25 stabilize the mixture, increase shelf life of the reagent mixture, reduce volatility, inhibit toxicity, inhibit photodegradation, and the like. Such additives are known to those skilled in the art and art not otherwise specified in detail herein.

Tables 1-13 following are examples of reagent mixtures usefully employed in the practice of the present invention for planarizing copper surfaces or other  
30 surfaces as indicated on the Tables.

TABLE 1: AQUEOUS PEROXIDE - PHOSPHORIC ACID REAGENT SOLUTIONS FOR PLANARIZATION OF COPPER

	<u>Oxidizer</u>	<u>Co-Reactant</u>	<u>Other Additives</u>
5	a) $\text{H}_2\text{O}_2$	$\text{H}_3\text{PO}_4$	HCl, aliphatic alcohols
	b) $\text{H}_2\text{O}_2$	$\text{H}_3\text{PO}_4$	HCl, Agidol (butylated hydroxytoluene)
10	c) $\text{H}_2\text{O}_2$	$\text{H}_3\text{PO}_4$	HCl, Agidol-2,
	d) $\text{H}_2\text{O}_2$	$\text{H}_3\text{PO}_4$	HCl, 2,6-di-tert-butyl- 4[(dimethylamino) methyl]phenol
15	e) $\text{H}_2\text{O}_2$	$\text{H}_3\text{PO}_4$	HCl; $\text{H}_3\text{PO}_4$ , $(\text{HPO}_4)^{2-}$ , $\text{PO}_4^{3-}$
	f) $\text{H}_2\text{O}_2$	$\text{H}_3\text{PO}_4$	HCl, 2,6-di-tert- -4N,N-dimethyl aminomethylphenol
20	g) $\text{H}_2\text{O}_2$	$\text{H}_3\text{PO}_4$	borax
	h) $\text{H}_2\text{O}_2$	$\text{H}_3\text{PO}_4$	various additives

TABLE 2: AQUEOUS PEROXIDE - SULFURIC ACID REAGENT SOLUTIONS FOR PLANARIZATION OF COPPER

	<u>Oxidizer</u>	<u>Co-Reactant</u>	<u>Other Additives</u>
5	a) $\text{H}_2\text{O}_2$	$\text{H}_2\text{SO}_4$ / $\text{H}_3\text{PO}_4$	Ethylene glycol, $\text{ZnSO}_4$
10	b) $\text{H}_2\text{O}_2$	$\text{H}_2\text{SO}_4$	MeOH, Poly(oxy ethylene)lauryl ether, Malic acid
	c) $\text{H}_2\text{O}_2$	$\text{H}_2\text{SO}_4$	$\text{HOOC}(\text{CX}_2)_n\text{COOH}$ with $\text{X}=\text{OH}$ , amine, H $n=1-4$
15	d) $\text{H}_2\text{O}_2$	$\text{H}_2\text{SO}_4$	3% tartaric acid 1% ethylene glycol
	e) $\text{H}_2\text{O}_2$	$\text{H}_2\text{SO}_4$	1,2,4-triazole, 1,2,3-triazole, tetrazole, nonionic surfactant
20	f) $\text{H}_2\text{O}_2$	$\text{H}_2\text{SO}_4$	$\text{C}_2\text{H}_5\text{OH}$ , aliphatic alcohols, nonionic surfactant
25	g) $\text{H}_2\text{O}_2$	$\text{H}_2\text{SO}_4$	Trifluoroethanol, Laprol 602 <sup>®</sup> surfactant, aliphatic alcohols
	h) $\text{H}_2\text{O}_2$	$\text{H}_2\text{SO}_4$	aliphatic alcohols
30	i) $\text{H}_2\text{O}_2$	$\text{H}_2\text{SO}_4$	$\text{SiF}_6$ , Organic salt surfactant
	j) $\text{H}_2\text{O}_2$	$\text{H}_2\text{SO}_4$	various additives

TABLE 3: AQUEOUS PEROXIDE MINERAL ACID REAGENT SOLUTIONS  
FOR PLANARIZATION OF COPPER

5	<u>Oxidizer</u>	<u>Co-Reactant</u>	<u>Other Additives</u>
10	a) $\text{H}_2\text{O}_2$	$\text{HNO}_3$	alcohols, $\text{HOOC}(\text{CX}_2)_n\text{COOH}$ $\text{X}=\text{OH}$ , amines, H $n=1-4$
	b) $\text{H}_2\text{O}_2$	$\text{HNO}_3$	various additives

TABLE 4: AQUEOUS NITRIC ACID REAGENT SOLUTIONS FOR  
PLANARIZATION OF COPPER

	<u>Oxidizer</u>	<u>Co-Reactant</u>	<u>Other Additives</u>
5	a) $\text{H}_2\text{O}_2$ / $\text{HNO}_3$	$\text{H}_3\text{PO}_4$	methanol
	b) $\text{H}_2\text{O}_2$ / $\text{HNO}_3$		Trifluoroethanol, Laprol 602 <sup>®</sup>
10			Surfactant, aliphatic alcohols
	c) $\text{HNO}_3$	$\text{H}_3\text{PO}_4$	PVA
	d) $\text{HNO}_3$	$\text{H}_2\text{SO}_4$	diphenylsulfamic acid, aliphatic alcohols
15	e) $\text{HNO}_3$	$\text{H}_2\text{SO}_4$	HCl
	f) $\text{HNO}_3$	$\text{H}_2\text{SO}_4$	various additives
	g) $\text{HNO}_3$		BTA (benzotriazole)
20			

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TABLE 5: AQUEOUS PEROXIDE ORGANIC ACID REAGENT SOLUTIONS  
FOR PLANARIZATION OF COPPER

	<u>Oxidizer</u>	<u>Co-Reactant</u>	<u>Other Additives</u>
5	a) $\text{H}_2\text{O}_2$	Oxalic acid	Sodium oxalate, Benzotriazole, Sodium Lignosulfonate
10	b) $\text{H}_2\text{O}_2$	other organic acids	various additives

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TABLE 6: AQUEOUS DILUTE MINERAL ACID REAGENT SOLUTIONS  
FOR PLANARIZATION OF COPPER

5	<u>Acid</u>	<u>Other Additives</u>
	a) $\text{H}_3\text{PO}_4$	various additives

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TABLE 7: AQUEOUS CONCENTRATED ACID REAGENT SOLUTIONS FOR  
PLANARIZATION OF COPPER

5				
		<u>Oxidizer</u>	<u>Acids</u>	<u>Other Additives</u>
	a)		H <sub>3</sub> PO <sub>4</sub> /Acetic/ H <sub>2</sub> SO <sub>4</sub>	
	b)		H <sub>3</sub> PO <sub>4</sub> /Acetic/ HNO <sub>3</sub>	
10	c)		H <sub>3</sub> PO <sub>4</sub> /Acetic/ HNO <sub>3</sub> / H <sub>2</sub> SO <sub>4</sub>	Glycol, Gelatine Carboxymethyl- cellulous, amines, surfactants, heavy metal salts including Cu and Ta.
15				
	d)	H <sub>2</sub> O <sub>2</sub>	H <sub>3</sub> PO <sub>4</sub> /Acetic/ H <sub>2</sub> SO <sub>4</sub>	Glycol, Gelatine Carboxymethyl- cellulous, amines, surfactants, heavy metal salts including Cu and Ta.
20		.		
	e)	H <sub>2</sub> O <sub>2</sub>	H <sub>3</sub> PO <sub>4</sub> /H <sub>2</sub> SO <sub>4</sub>	100ml propylene glycol, 100ml 2-ethyl- hexylamine, 25 ppm Cl <sup>-</sup> .
25				
	f)		H <sub>3</sub> PO <sub>4</sub> /Acetic/ HNO <sub>3</sub>	nonionic surfactant
30				
	g)	H <sub>2</sub> O <sub>2</sub>	H <sub>3</sub> PO <sub>4</sub> /Acetic/ HNO <sub>3</sub> / H <sub>2</sub> SO <sub>4</sub>	various additives
	h)		H <sub>3</sub> PO <sub>4</sub> /HNO <sub>3</sub>	

TABLE 8: AQUEOUS DILUTE ACID - METAL SALT REAGENT SOLUTIONS FOR PLANARIZATION OF COPPER

5

	<u>Oxidizer</u>	<u>Acid</u>	<u>Metal Salt</u>	<u>Other Additives</u>
	a)	HCl	CuCl	
	b)	HCl	CuCl	KCl
	c)	HCl	FeCl	various additives
10	d) H <sub>2</sub> O <sub>2</sub>	H <sub>2</sub> SO <sub>4</sub>	CuCl	n-propanol
	e)	HCl	CuCl	various additives
	f) H <sub>2</sub> O <sub>2</sub>	H <sub>2</sub> SO <sub>4</sub>	CuCl	various additives
	g)	HCl	FeCl <sub>3</sub>	glycerol
	h) HNO <sub>3</sub>	HCl	FeCl <sub>3</sub>	
15	i)	HCl	FeCl <sub>3</sub>	
	j)	HCl	FeCl <sub>3</sub>	various additives
	k)	HCl	FeCl <sub>3</sub>	CuCl <sub>2</sub> , SnCl <sub>2</sub>
	l)	HCl	FeCl <sub>3</sub>	ethanol

TABLE 9: AQUEOUS OXIDIZER - SALT REAGENT SOLUTIONS FOR PLANARIZATION OF COPPER

	<u>Oxidizer</u>	<u>2nd Oxidizer</u>	<u>Base</u>	<u>Salt</u>	<u>Other Additives</u>
5	a) $\text{NaClO}_3$		$\text{NH}_4\text{F}$	$\text{CuSO}_4$	Na EDTA salt of wetting agent
	b) $\text{FeNO}_3$				various additives
	c) $(\text{NH}_4)_2\text{S}_2\text{O}_8$				various additives
10	d) $\text{CuNH}_4\text{Cl}_3$		$\text{NH}_4\text{OH}$		various additives
	e) $\text{Na}_2\text{S}_2\text{O}_3$	$\text{K}_2\text{S}_2\text{O}_5$			various additives

TABLE 10: AQUEOUS BASE REAGENT SOLUTIONS FOR  
PLANARIZATION OF COPPER

5	<u>Base</u>	<u>Oxidizer</u>	<u>Other Additives</u>
	a) $\text{NH}_4\text{OH/KOH}$	$\text{H}_2\text{O}_2$	various additives
	b) $\text{NH}_4\text{OH}$	$\text{H}_2\text{O}_2$	various additives
	c) $\text{NH}_4\text{OH}$	$(\text{NH}_4)_2\text{S}_2\text{O}_8$	various additives
	d) $\text{NH}_4\text{OH}$		$\text{Cu}(\text{NO}_3)_2$

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TABLE 11: AQUEOUS ACID REAGENT SOLUTIONS FOR  
PLANARIZATION OF TANTALUM AND COPPER

5	<u>Oxidizer</u>	<u>Acid</u>	<u>Other Additives</u>
	a) $\text{HNO}_3$	HF	various additives
	b) $\text{H}_2\text{O}_2$	HF	various additives
	c) $\text{HNO}_3$	HF	lactic acid, various additives
	d) $\text{H}_2\text{O}_2$	HF	$\text{H}_2\text{SO}_4$

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TABLE 12: AQUEOUS BASE REAGENT SOLUTIONS FOR  
PLANARIZATION OF TANTALUM AND COPPER

	<u>Base</u>	<u>Acid</u>	<u>Other Additives</u>
5	a) NaOH		
	b) NaOH	H <sub>2</sub> O <sub>2</sub>	
	c) KOH	H <sub>2</sub> O <sub>2</sub>	
	d) NH <sub>4</sub> OH	H <sub>2</sub> O <sub>2</sub>	
10	Remove oxide film after CMP with rinse of dilute acetic acid, dilute nitric acid, aqueous solution or other solutions.		

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TABLE 13: MISCELLANEOUS REAGENTS FOR PLANARIZATIONS  
OF COPPER

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- a) EDTA,  $\text{NH}_4\text{OH}$ ,  $\text{H}_2\text{O}_2$ , in aqueous solution
- 5 b) Citric acid, Erythorbic acid, Triethanolamine, in aqueous solution
- c) Trisodium citrate, Triethanolamine, Sodium nitrate, in aqueous solution
- d)  $\text{H}_2\text{SO}_4$ ,  $\text{H}_2\text{O}_2$ , Sodium molybdate, Phenolsulfonic acid, in aqueous solution
- e) Mineral acid (sulfuric,  $\text{HCl}$  or the like), molybdenum salt
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In addition to the additives shown in Tables 1-13 above, other additives include but are not limited to the following:

- borax, zinc sulfate, copper carbonate, alcohol (including low molecular weight alcohols, glycols, phenols, aliphatic alcohols, polyvinylalcohols and the
- 15 like), surfactants (including anionic, cationic, nonionic surfactants and others preferentially adhering to certain materials, modifying thereby the chemical reactivity of certain sites), solution stabilizers (including polyvinyl alcohols and other agents inhibiting spontaneous decomposition of oxidizing agents), wetting agents.

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For example, one specific formulation uses highly concentrated  $H_2O_2$  in a strong base, such as KOH, plus chemical-active abrasive particles such as  $CeO_2$  or  $SnO_2$ , in conjunction with other chemical components. According to Tytgat (4,981,553), the chemical etchant alone (without abrasion) provides a Ta removal rate of approximately 1000-2000 Ångstrom/min., which is significantly higher than Ta removal rates typically available Cu slurries. The present invention may be expected to achieve increased Ta removal rates in CMP by employing the additional mechanism of mechanical abrasion of the wafer by abrasive particles. In addition, heating the slurry, or the opposing body or pad, could enhance the removal rate of the Ta or TaN layers. Thus temperature control may be another means of achieving a low polishing pressure CMP process.

Having described the invention in detail, those skilled in the art will appreciate that, given the present disclosure, modifications may be made to the invention without departing from the spirit of the inventive concept described herein. Therefore, it is not intended that the scope of the invention be limited to the specific and preferred embodiments illustrated and described. Rather, it is intended that the scope of the invention be determined by the appended claims.

CLAIMS

We Claim:

- 5 1) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface comprising:
  - a) an oxidizing reactant selected from the group consisting of  $\text{H}_2\text{O}_2$ ,  $\text{HNO}_3$  and mixtures thereof; and,
  - b) a co-reactant is selected from the group consisting of  $\text{H}_3\text{PO}_4$ ,  $\text{H}_2\text{SO}_4$ ,  
 10  $\text{HNO}_3$ , oxalic acid, acetic acid, organic acids and mixtures thereof; and,
  - c) other additives selected from the group consisting of selected from the group consisting of HCl, aliphatic alcohols, butylated hydroxytoluene, Agidol-2, 2,6-di-tert-butyl-4[(dimethylamino)methyl]phenol, 2,6-di-tert-4N,N-dimethylaminomethylphenol, borax, ethylene glycol,  $\text{ZnSO}_4$ , methanol, propanol,  
 15 poly(oxyethylene)lauryl ether, malic acid,  $\text{HOOC}(\text{CX}_2)_n\text{COOH}$  wherein  $\text{X}=\text{OH}$ , amine, H and  $n=1-4$ ), 3% tartaric acid, 1% ethylene glycol, 1,2,4-triazole, 1,2,3-triazole, tetrazole, nonionic surfactant, ethanol, trifluoroethanol,  $\text{SiF}_6$ , organic salt surfactant, polyvinyl alcohol, diphenylsulfamic acid, sodium oxalate, benzotriazole, sodium lignosulfonate, glycol, gelatin carboxymethylcellulose,  
 20 amines, heavy metal salts, salts of Cu and Ta, KCl,  $\text{CuCl}_2$ ,  $\text{SnCl}_2$ , propylene glycol, 2-ethyl-hexylamine, copper carbonate, low molecular weight alcohols, glycols, phenols, aliphatic alcohols, polyvinylalcohols, anionic surfactants, cationic surfactants, fluorocarbon-based surfactants, nonionic surfactants having the properties of preferentially adhering to certain materials, modifying thereby the  
 25 chemical reactivity where so adhered, polyvinyl alcohol solution stabilizers and species inhibiting spontaneous decomposition of oxidizing agents, wetting agents and mixtures thereof.
- 2) An etching solution as in claim 1 further comprising a species selected from  
 30 the group consisting of  $\text{CuCl}$ ,  $\text{FeCl}$ ,  $\text{FeCl}_3$ , and mixtures thereof.

- 3) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface comprising species selected from the group consisting of  $\text{NaClO}_3$ ,  $\text{FeNO}_3$ ,  $(\text{NH}_4)_2\text{S}_2\text{O}_8$ ,  $\text{CuNH}_4\text{Cl}_3$ ,  $\text{Na}_2\text{S}_2\text{O}_8$ ,  $\text{K}_2\text{S}_2\text{O}_5$ ,  $\text{NH}_4\text{F}$ ,  $\text{CuSO}_4$ ,  $\text{NH}_4\text{OH}$ , sodium EDTA salt of wetting agent and mixtures thereof.
- 4) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface comprising species selected from the group consisting of,  $(\text{NH}_4)_2\text{S}_2\text{O}_8$ ,  $\text{KOH}$ ,  $\text{NH}_4\text{OH}$ ,  $\text{H}_2\text{O}_2$ ,  $\text{Cu}(\text{NO}_3)_2$  and mixtures thereof.
- 5) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface comprising species selected from the group consisting of  $\text{HF}$ ,  $\text{HNO}_3$ ,  $\text{H}_2\text{O}_2$ ,  $\text{H}_2\text{SO}_4$ , lactic acid and mixtures thereof.
- 6) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface comprising species selected from the group consisting of,  $\text{NaOH}$ ,  $\text{KOH}$ ,  $\text{NH}_4\text{OH}$ ,  $\text{H}_2\text{O}_2$ , and mixtures thereof.
- 7) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface comprising: EDTA,  $\text{NH}_4\text{OH}$ ,  $\text{H}_2\text{O}_2$ , in aqueous solution.
- 8) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface comprising: citric acid, erythorbic acid, triethanolamine, in aqueous solution.
- 9) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface comprising: trisodium citrate, triethanolamine, sodium nitrate, in aqueous solution.

- 10) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface comprising:  $\text{H}_2\text{SO}_4$ ,  $\text{H}_2\text{O}_2$ , sodium molybdate, phenolsulfonic acid, in aqueous solution.
- 5 11) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface comprising: mineral acid, molybdenum salt.
- 12) An etching solution for chemical mechanical planarization of a Cu/Ta/TaN surface as in claim 1 further comprising abrasive particles selected from the group
- 10 consisting  $\text{SiO}_2$ ,  $\text{Al}_2\text{O}_3$  metallic and solid elemental particles, polymer particles, oxides, carbides, fluorides, carbonates, borides, nitrides, hydroxides of Al, Ag, Au, Ca, Ce, Cr, Cu, Fe, Gd, Ge, La, In, Hf, Mn, Mg, Ni, Nd, Pb, Pt, P, Sb, Sc, Sn, Tb, Ti, Ta, Th, Y, W, Zn, Zr, and mixtures thereof.
- 15 13) An etching solution as in claim 12 wherein said abrasive particles are coated.
- 14) An etching solution as in claim 13 wherein said coating is a chemically active species.
- 20 15) An etching solution as in claim 12 wherein said coating is  $\text{CeO}_2$ .
- 16) An etching solution as in claim 12 wherein said particles are produced by the sol method.
- 25 17) An etching solution as in claim 12 wherein said particles have a range of sizes from approximately 4 nanometers to approximately 5 micrometers.
- 18) An etching solution as in claim 12 wherein said particles have a size less than approximately 5 micrometers.

ABSTRACT

Chemical mechanical planarization of surfaces of copper, tantalum and tantalum  
5 nitride is accomplished by means of the chemical formulations of the present  
invention. The chemical formulations may optionally include abrasive particles  
and which may be chemically reactive or inert. Contact or noncontact CMP may  
be performed with the present chemical formulations. Substantially 1:1 removal  
rate selectivity for Cu and Ta/TaN is achieved.

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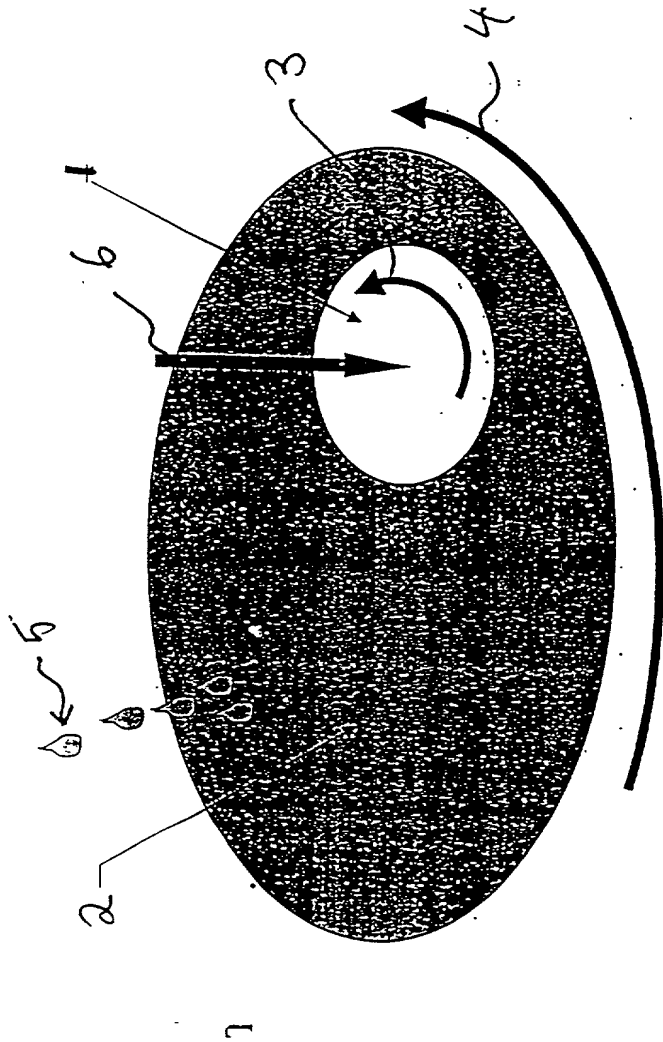


FIGURE 1

## DECLARATION FOR PATENT APPLICATION AND POWER OF ATTORNEY

As a below named inventor, I hereby declare that:

My residence, post office address and citizenship are as stated below adjacent to my name.

I believe I am the original, first and sole inventor (if only one name is listed below) or an original, first and joint inventor (if plural names are listed below) of subject matter (process, machine, manufacture, or composition of matter, or an improvement thereof) which is claimed and for which a patent is sought by way of the application entitled

### Composition for Chemical Mechanical Planarization of Copper, Tantalum and Tantalum Nitride

which (check) ☒ is attached hereto.

☐ and is amended by the Preliminary Amendment attached hereto.

☐ was filed on \_\_\_\_\_ as Application Serial No. \_\_\_\_\_

☐ and was amended on \_\_\_\_ (if applicable).

I hereby state that I have reviewed and understand the contents of the above identified specification, including the claims, as amended by any amendment referred to above.

I acknowledge the duty to disclose information, which is material to patentability as defined in Title 37, Code of Federal Regulations, § 1.56.

I hereby claim foreign priority benefits under Title 35, United States Code, § 119(a)-(d) of any foreign application(s) for patent or inventor's certificate or any PCT international application(s) designating at least one country other than the United States of America listed below and have also identified below any foreign application(s) for patent or inventor's certificate or any PCT international application(s) designating at least one country other than the United States of America filed by me on the same subject matter having a filing date before that of the application(s) of which priority is claimed:

Prior Foreign Application(s)			Priority Claimed	
Number	Country	Day/Month/Year Filed	Yes	No
N/A			<input type="checkbox"/>	<input type="checkbox"/>

I hereby claim the benefit under Title 35, United States Code, § 119(e) of any United States provisional application(s) listed below:

Provisional Application Number	Filing Date
N/A	

I hereby claim the benefit under Title 35, United States Code, § 120 of any United States application(s) or PCT international application(s) designating the United States of America listed below and, insofar as the subject matter of each of the claims of this application is not disclosed in the prior application(s) in the manner provided by the first paragraph of Title 35, United States Code, § 112, I acknowledge the duty to disclose information, which is material to patentability as defined in Title 37, Code of Federal Regulations, § 1.56, which became available between the filing date of the prior application(s) and the national or PCT international filing date of this application:

Application Serial No.	Filing Date	Status (patented, pending, abandoned)
N/A		

I hereby appoint the following attorney(s) and/or agent(s) to prosecute this application and to transact all business in the United States Patent and Trademark Office connected therewith:


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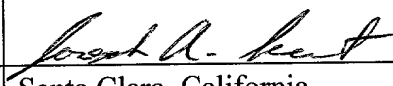
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